

This quiz is take-home and open book, and it is intended that all members of the group contribute to completing it. It is a violation of the Academic Honor Code to sign a quiz that you did not work on. **The quiz is due at the end of class on Thursday, January 28.**

List names in alphabetical order, and give social security numbers! Put names on all pages, and stapel pages together

Points

- (2) 1. Give the proper structural formulas for the following ionic compounds:
- (a) sodium sulfide (b) magnesium oxide
- (c) potassium selenide (d) calcium fluoride
- (3) 2. Draw Lewis dot structures for the following molecular compounds
- (a) NCl_3 (N is central atom) (b) CS_2 (C is central atom)
- (c) HOCl (O is central atom) (d) Chloroform (CHCl_3 , C is central atom)
- (e) Formaldehyde (CH_2O , C is central) (e) Hydrogen cyanide (HCN , C is central)

