This quiz is take-home and open book, and it is intended that all members of the group contribute to completing it. It is a violation of the Academic Honor Code to sign a quiz that you did not work on. **The quiz is due at the end of class on Thursday, February 18.**

List names in alphabetical order, and give social security numbers! Put names on all pages, and stapel pages together

_____   _____

_____   _____

_____   _____

Points

1.   Given that the energy produced from combustion of **one mole** of the following substances is as follows:

Substance combusted	Energy produced (kJ/mole)
Kerosene ($C_{14}H_{30}$)	8700
Propane (C_3H_8)	2,000
Ethanol (C_2H_5OH)	1,250
Glucose ($C_6H_{12}O_6$)	2,800
Coal (C)	360

(5)   (a) Calculate the energy produced by burning **one pound** of each of these substances. (One pound = 455 grams)

1. (con't.)

(10) (b) Calculate the **grams of CO$_2$** produced from each of these substances if one burned enough of the substance to generate 10^6 kJ of energy. (You will have to write the combustion reaction for each in order to work this one.)